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15: Chemical Equilibrium. We introduced the concept of equilibrium in Chapter 11, where you learned that a liquid and a vapor are in equilibrium when the number of molecules evaporating from the surface of the liquid per unit time is the same as the number of molecules condensing from the vapor phase. Vapor pressure is an example of a physical equilibrium because only the physical form of the substance changes.

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Chapter 15 Chemical Equilibrium SY 3/16/11 15-3 As the reaction proceeds, the product concentration increases and the reverse reaction begins to take place. $\text{FeSCN}_2^+(\text{aq}) \rightleftharpoons \text{Fe}^{3+}(\text{aq}) +$

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SCN-(aq) Rate (reverse) = $k_{\text{reverse}} [\text{FeSCN}^{2+}]$ As the reaction continues, the rate of the forward reaction decreases (because $[\text{Fe}^{3+}]$ and

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Chapter 13; Chapter 14; Chapter 15; Chapter 16; Chapter 17; Chapter 18; Chapter 19; Chapter 20; Chapter 21; Index; Learning Objectives. By the end of this section, you will be able to: Describe the nature of equilibrium systems; Explain the dynamic nature of a chemical equilibrium;

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